

CHIMIE - PHYSIQUE

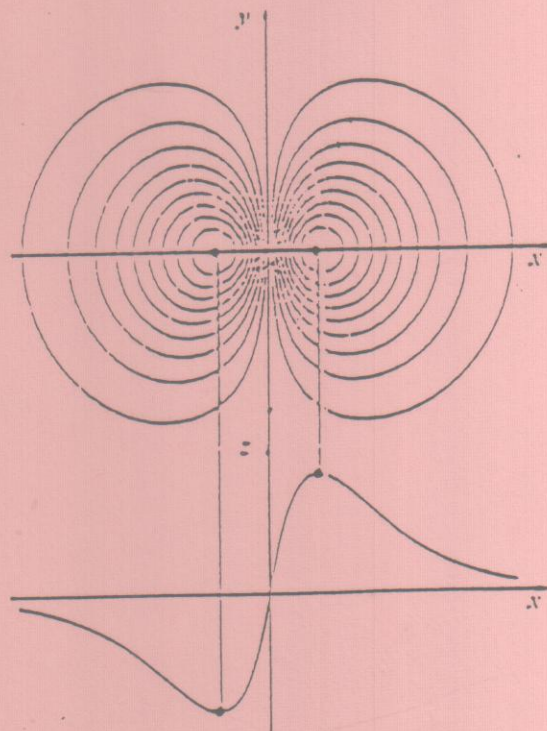
STRUCTURE

DE LA MATIÈRE

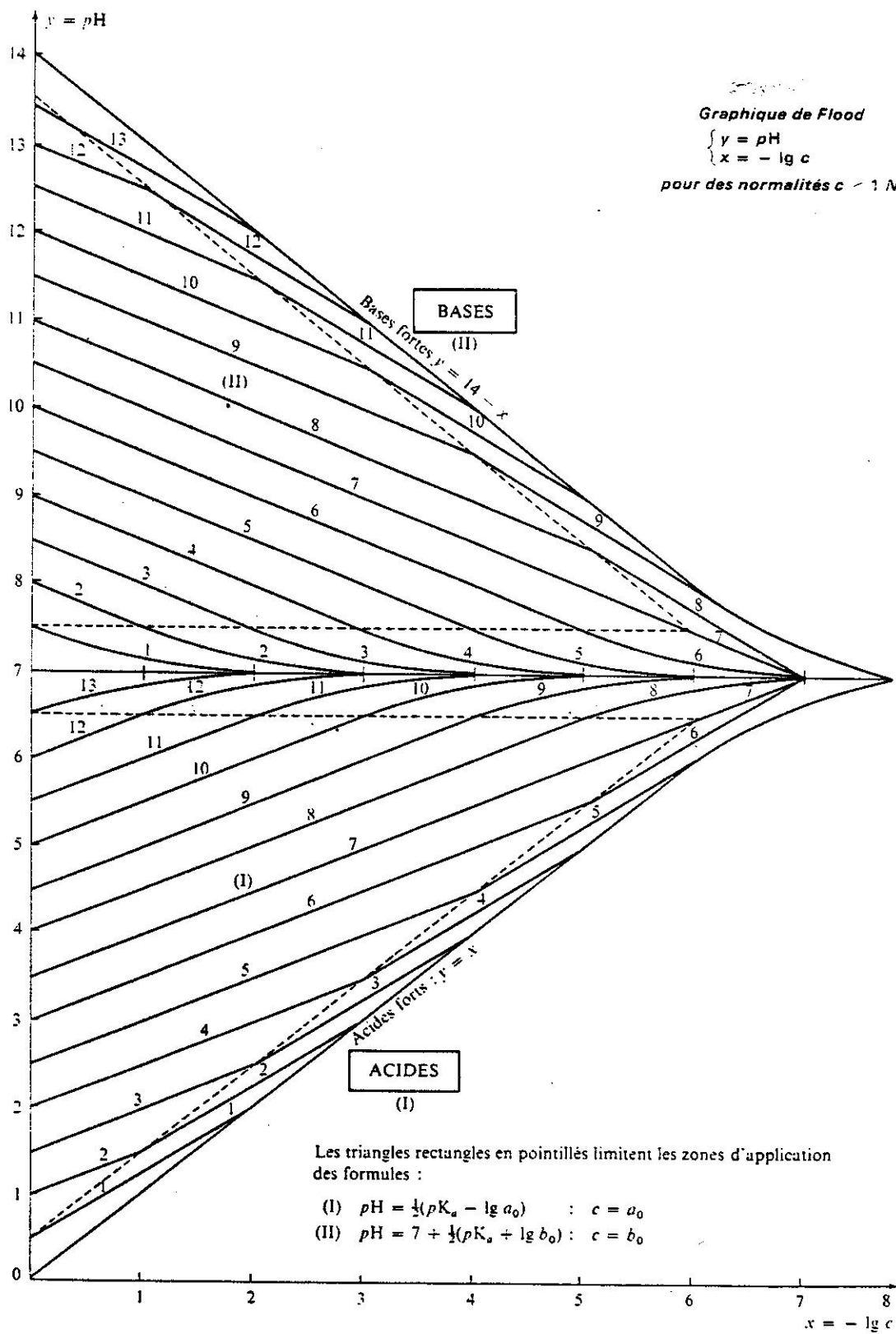
FICHES DU COURS

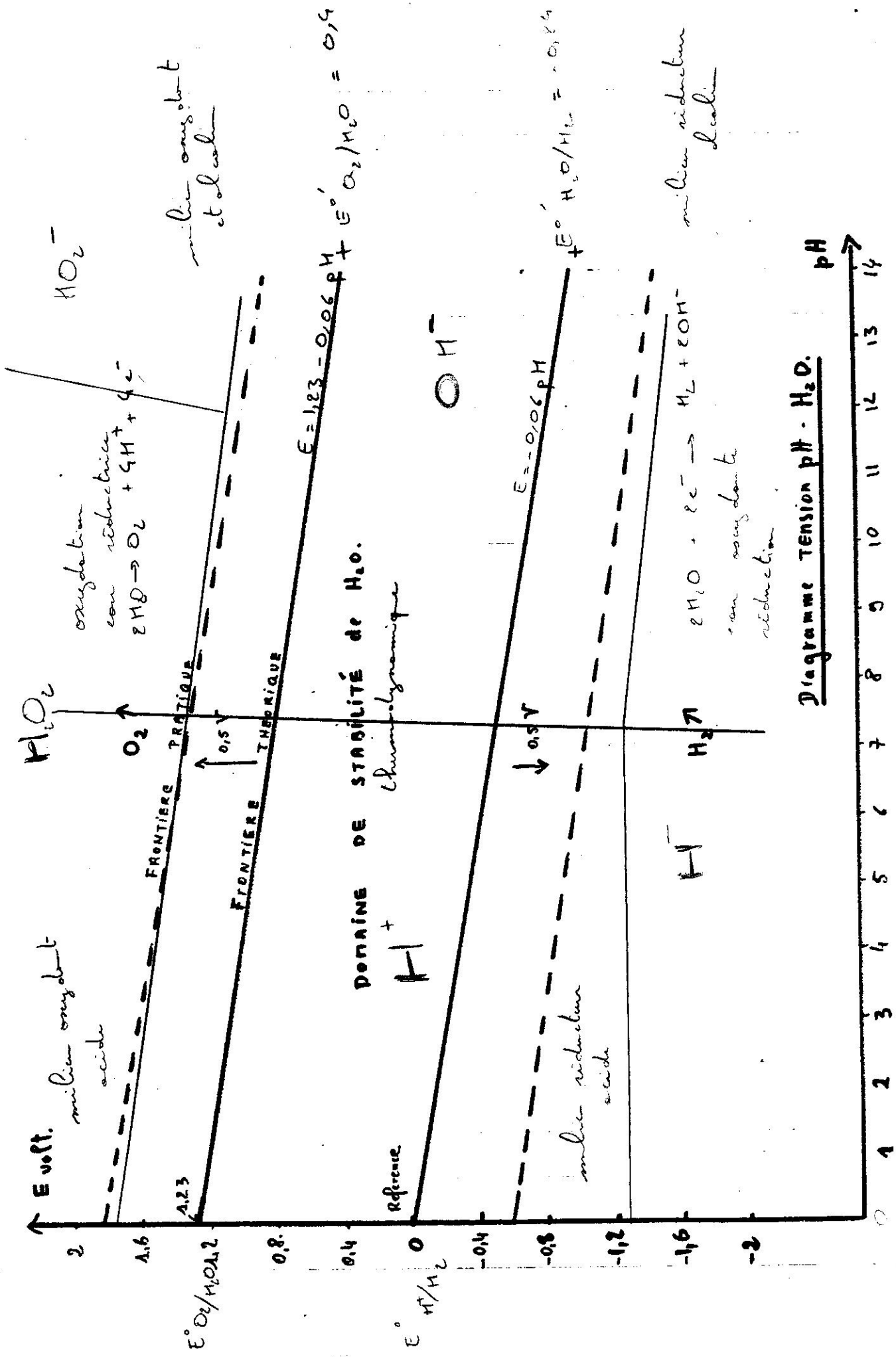
1 ère ANNEE . 1991-1992 .

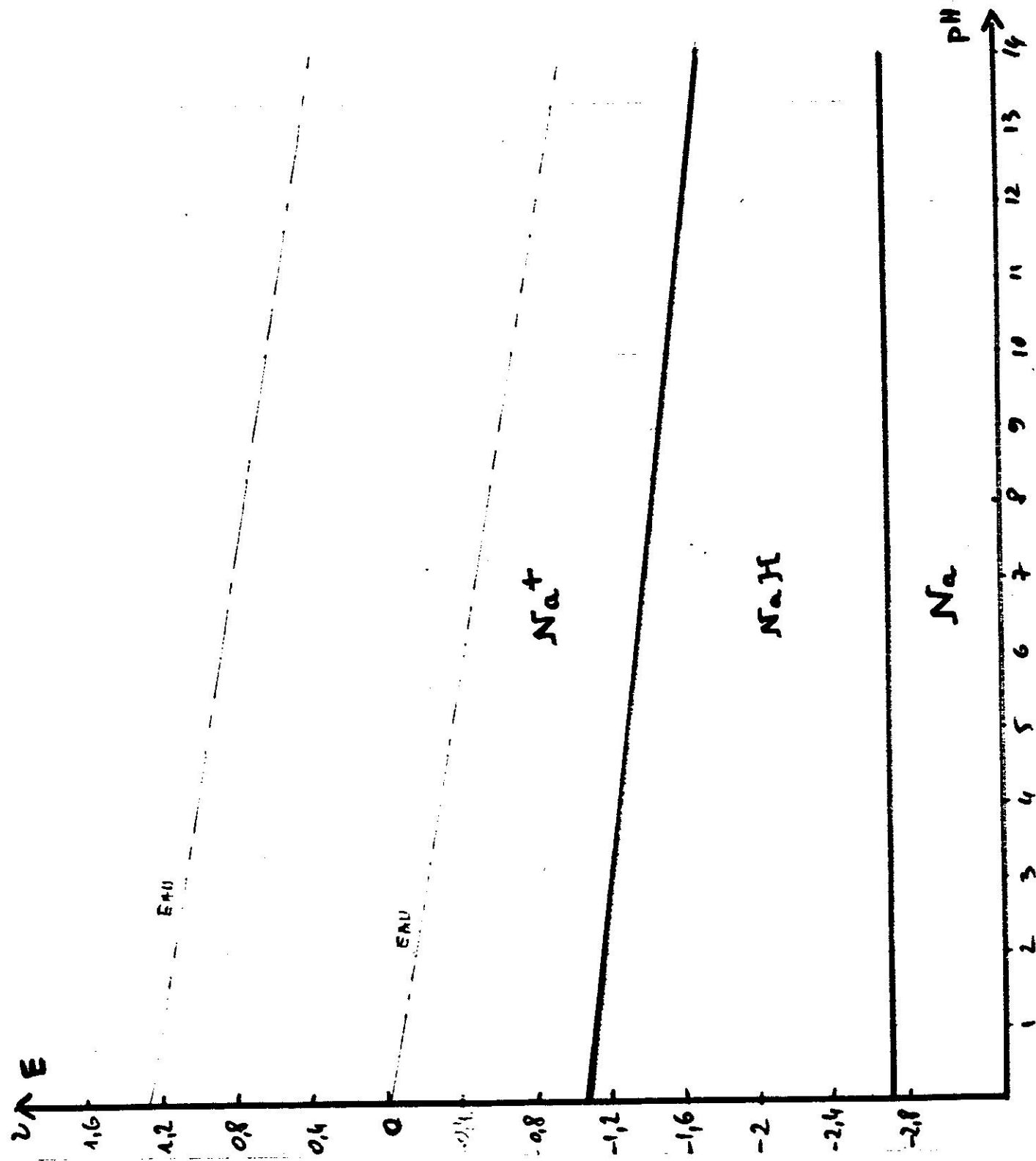
Pr. H. PINATEL

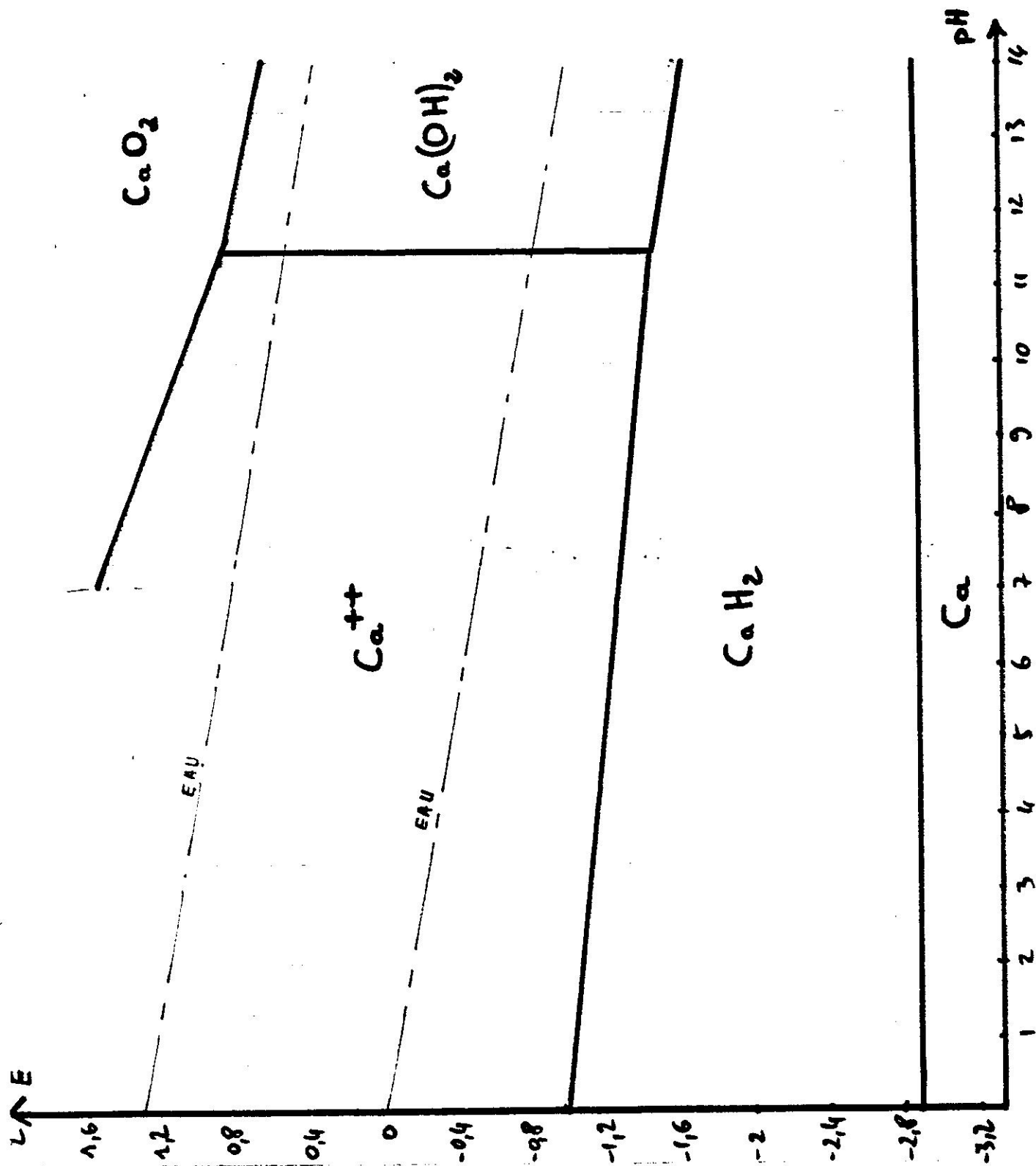


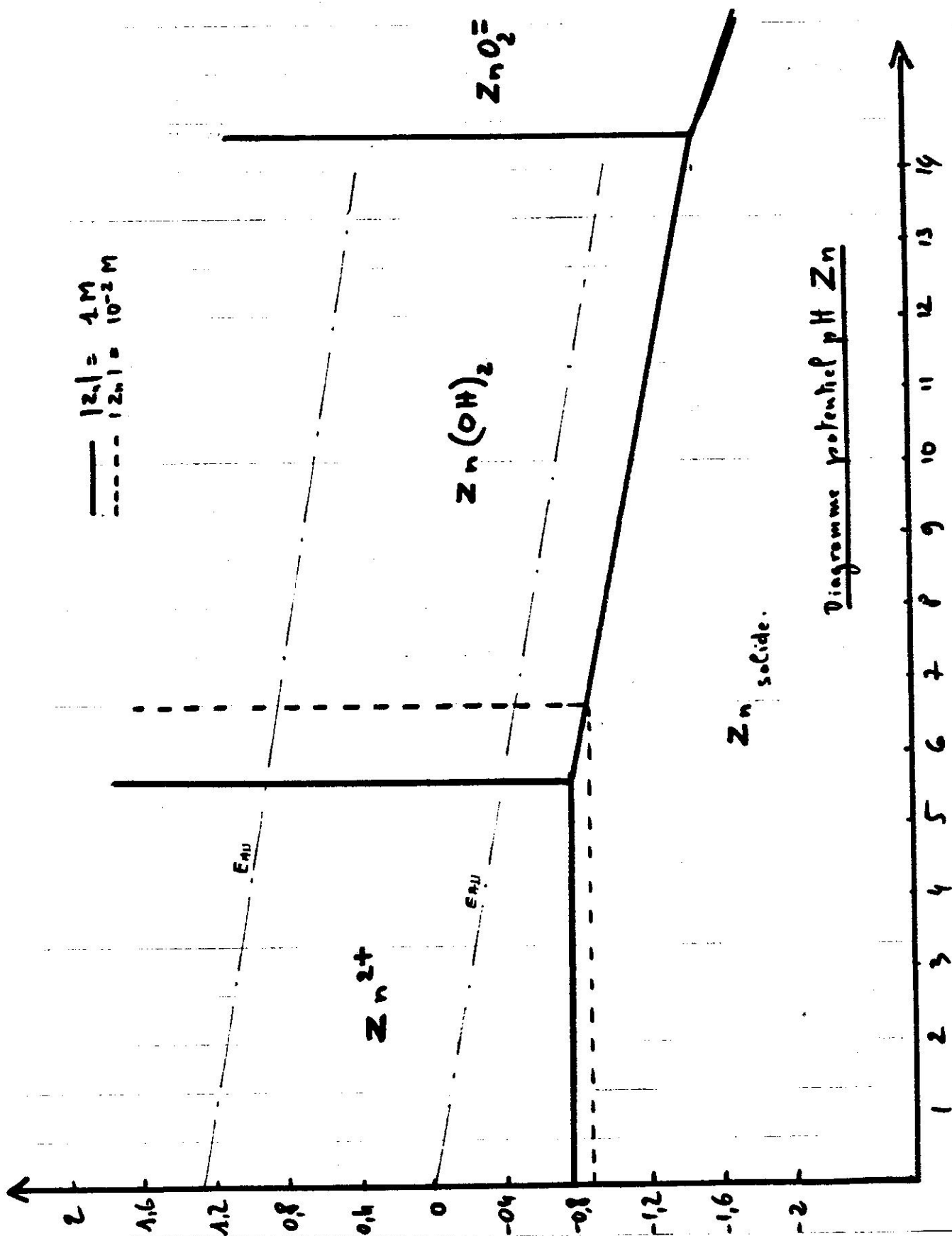
2ÈME PARTIE

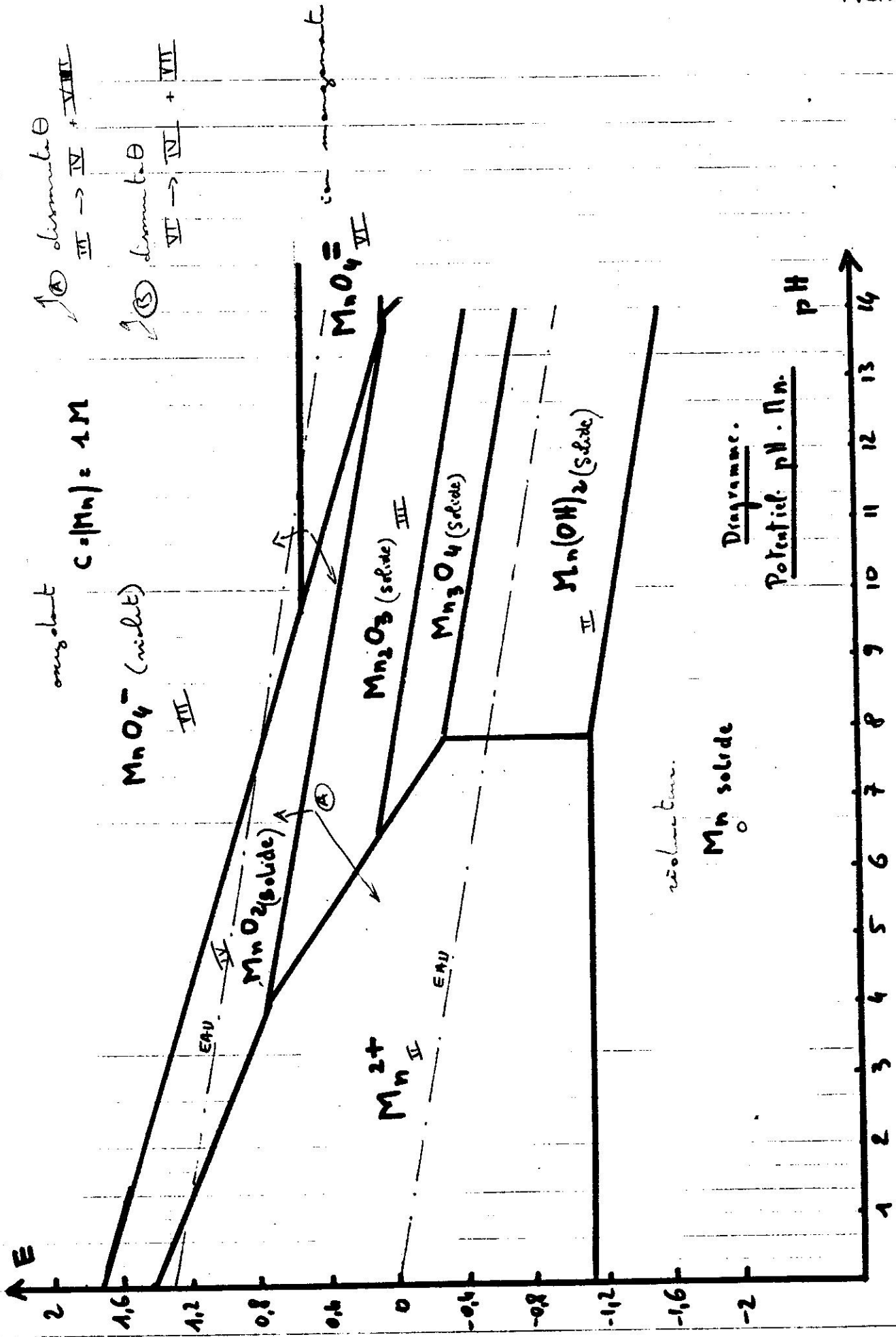


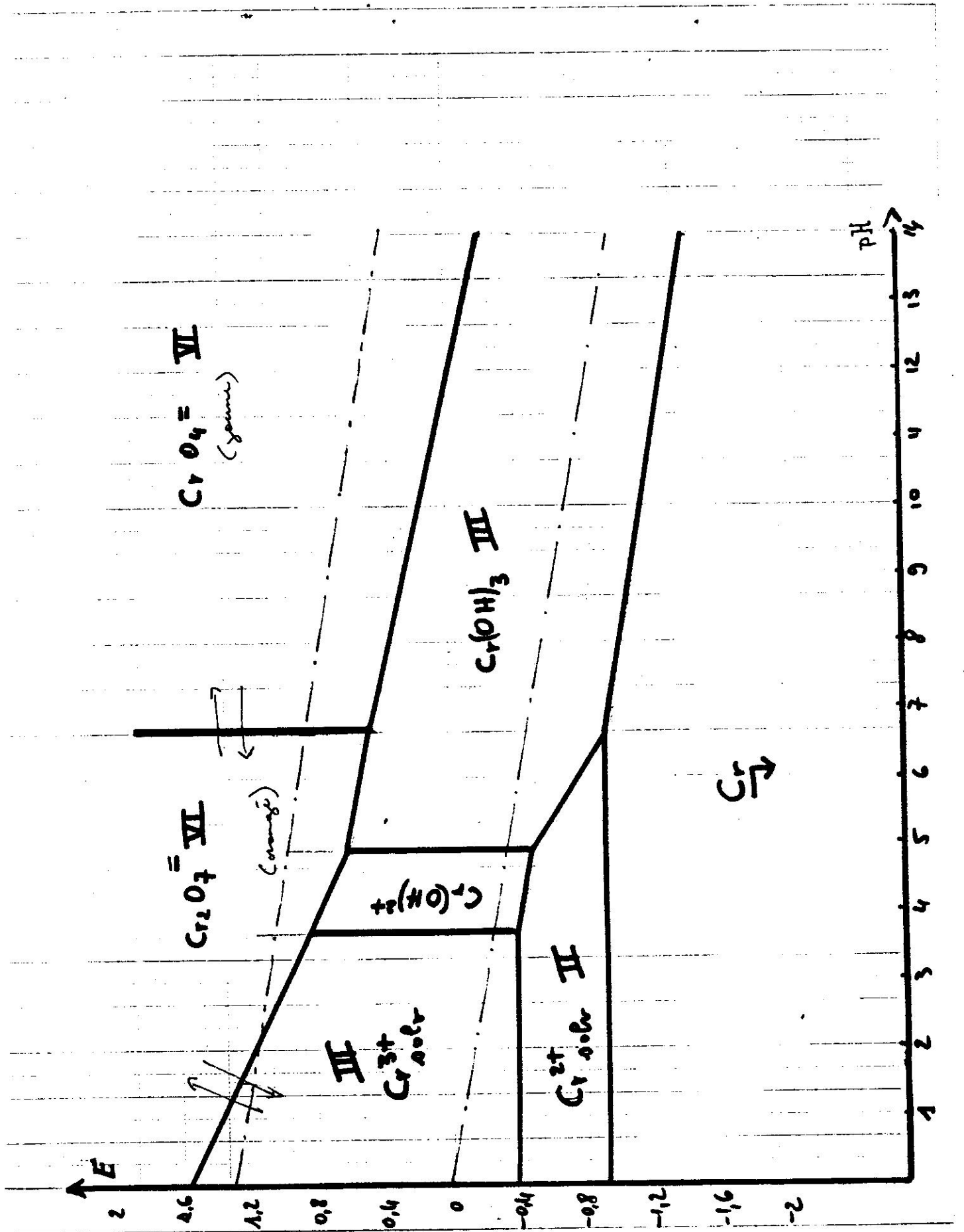


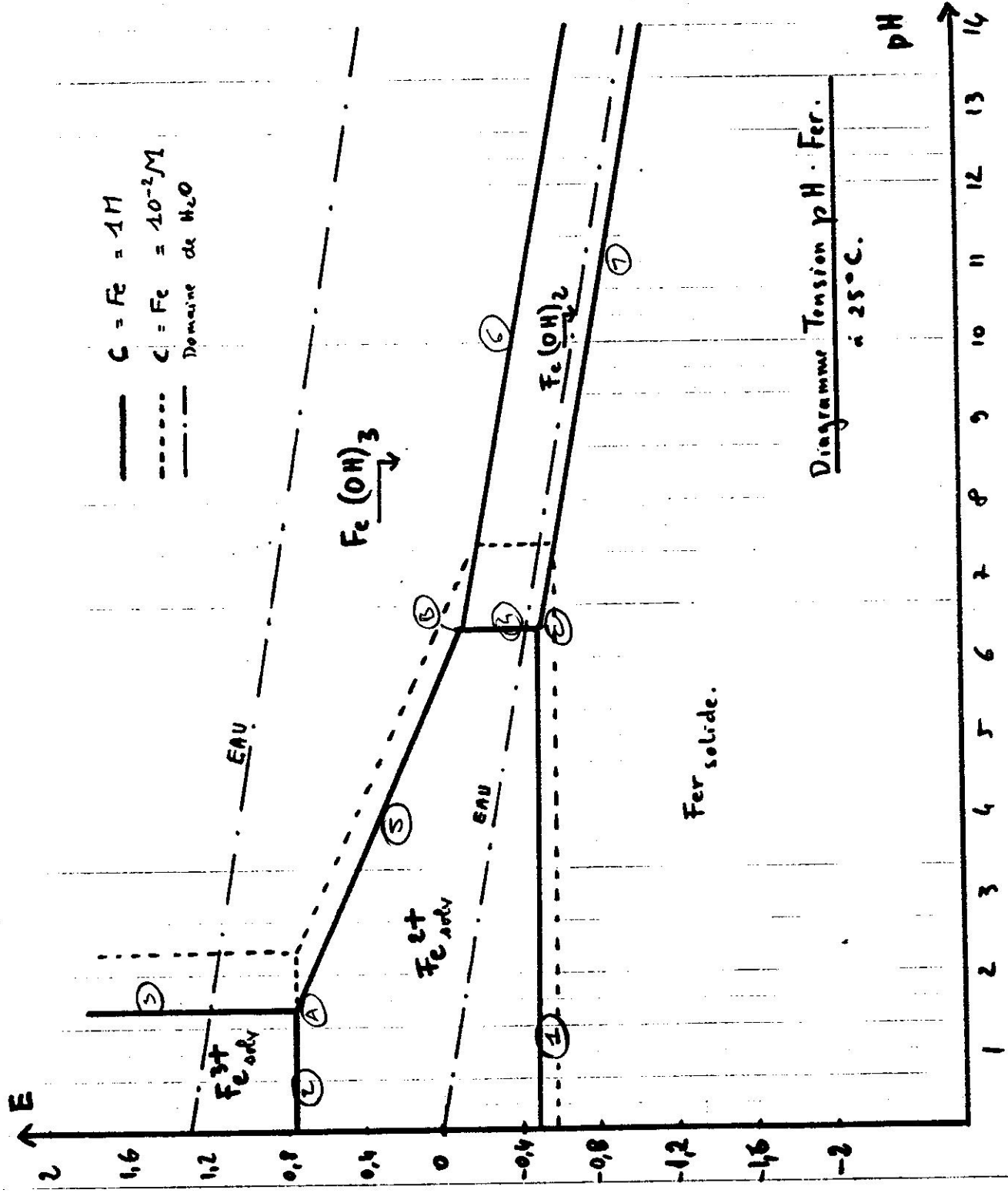










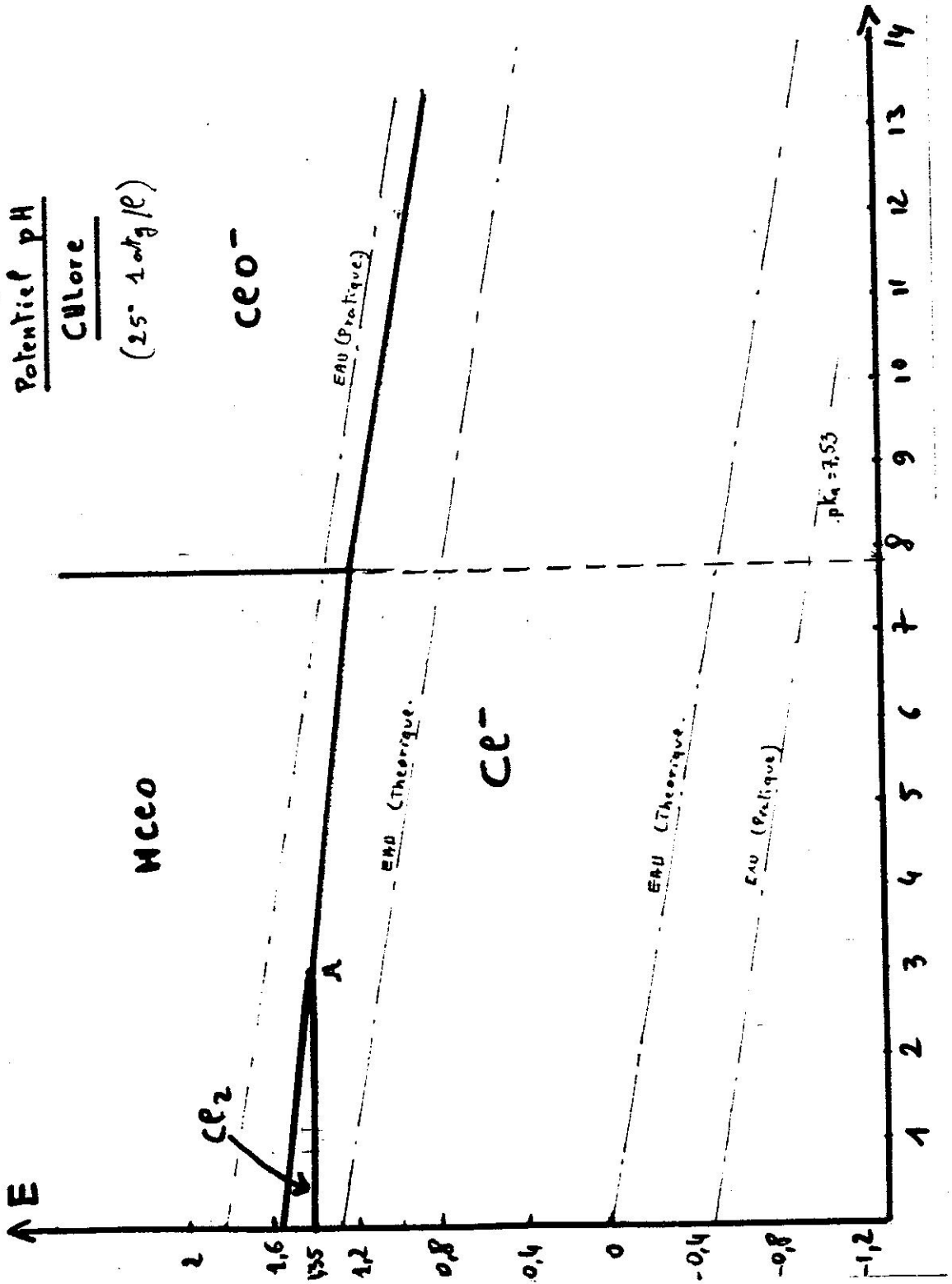


Diagramme

Potentiel pH

CU Lore

(25° 1.0g/l)



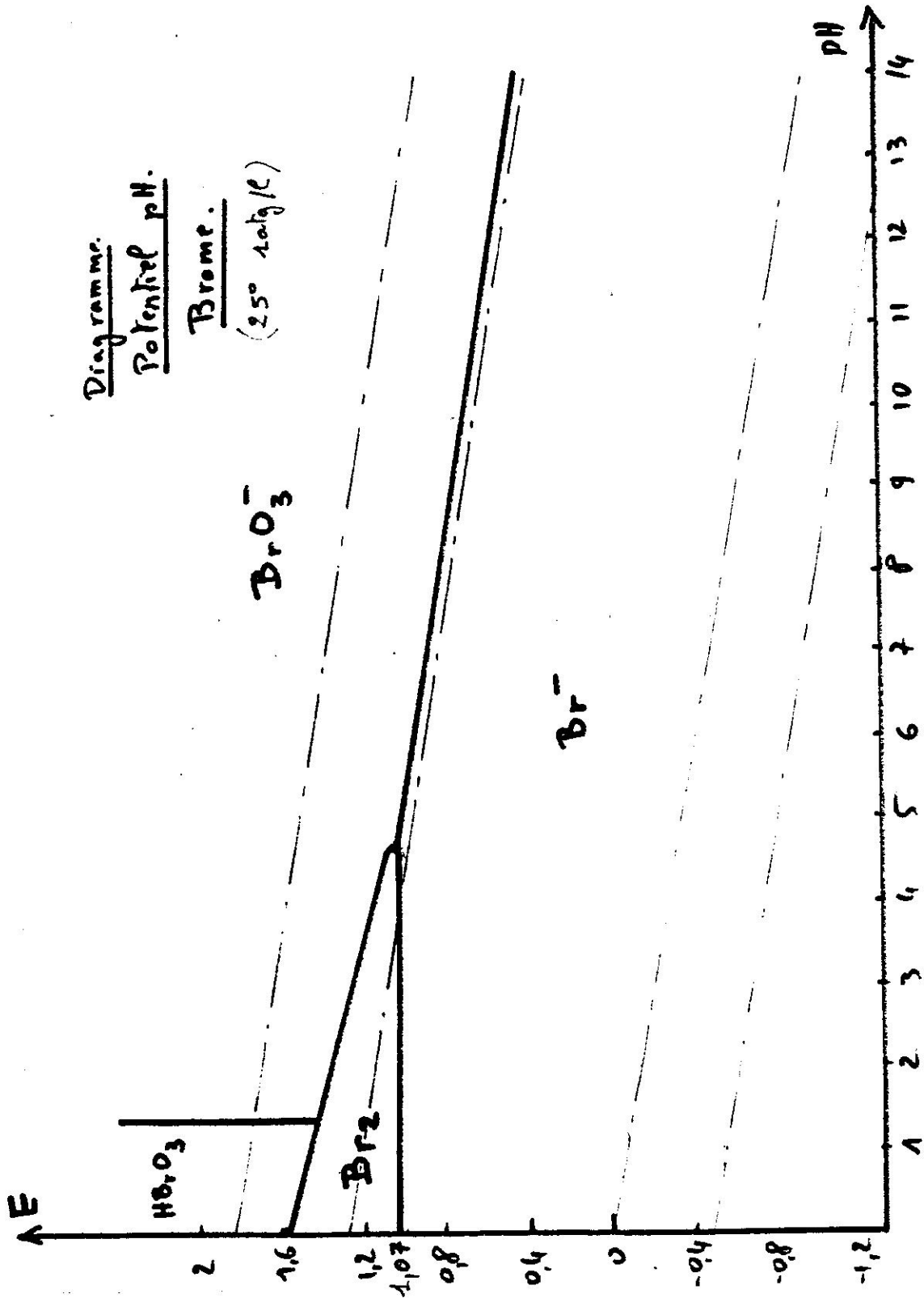
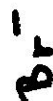
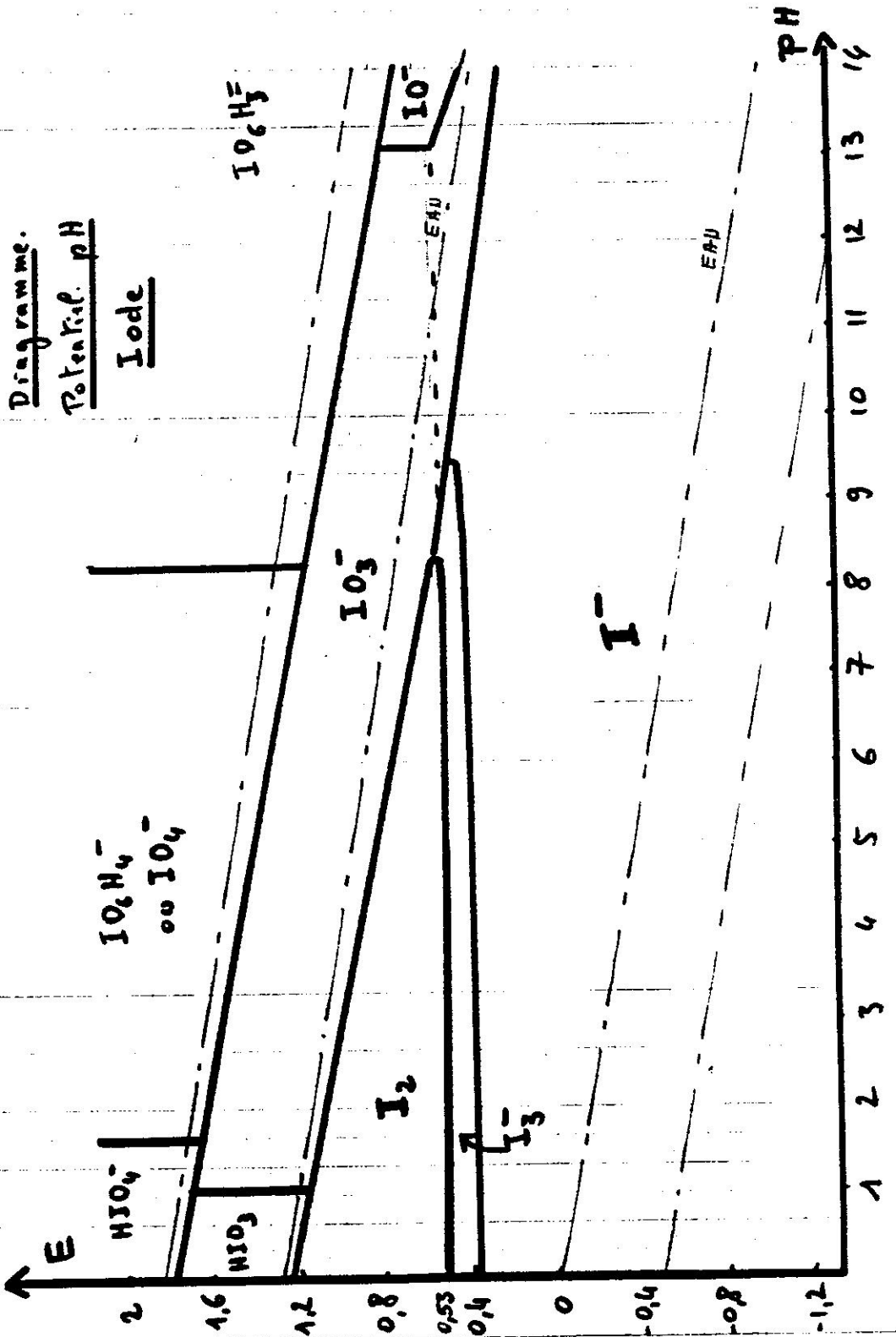
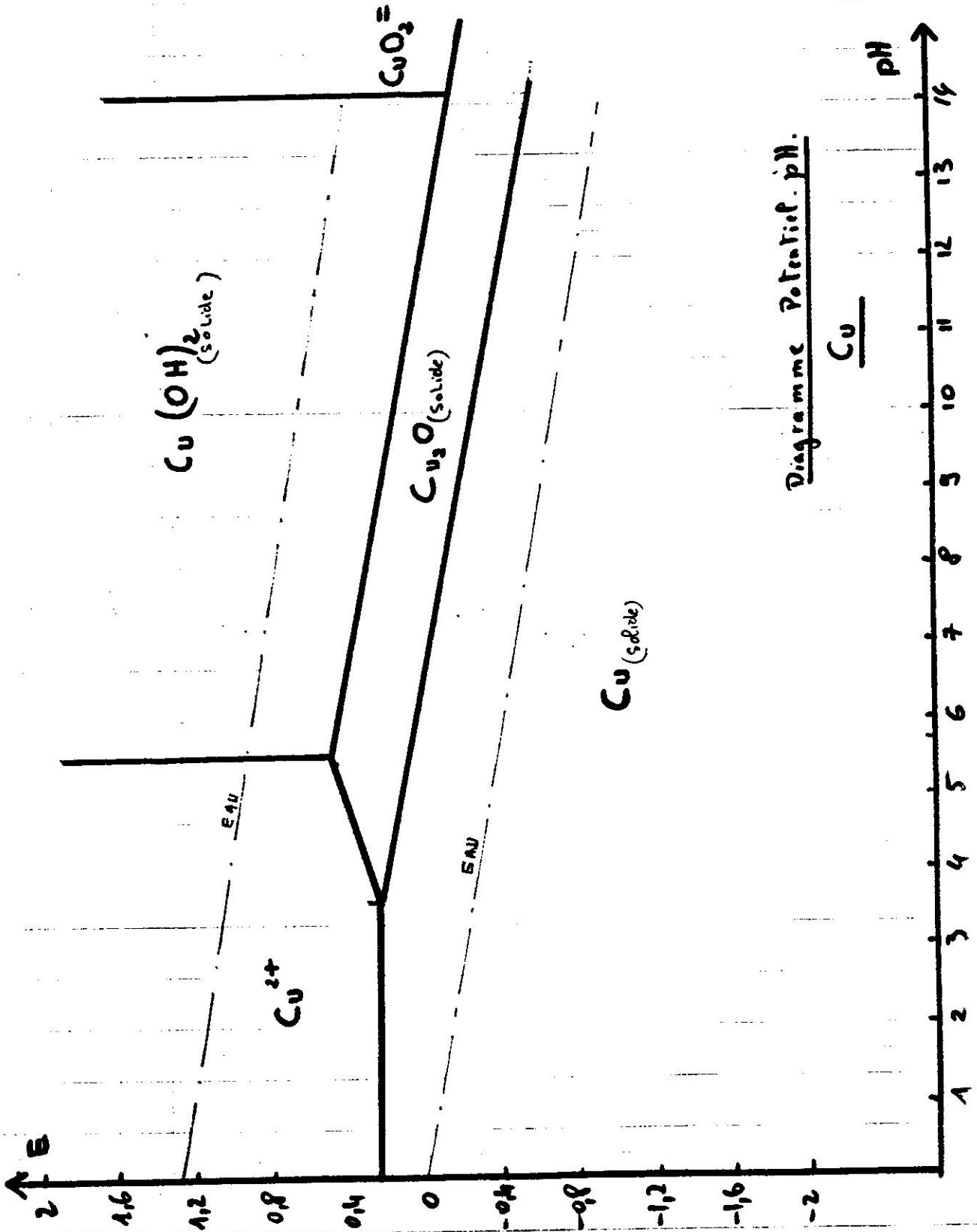
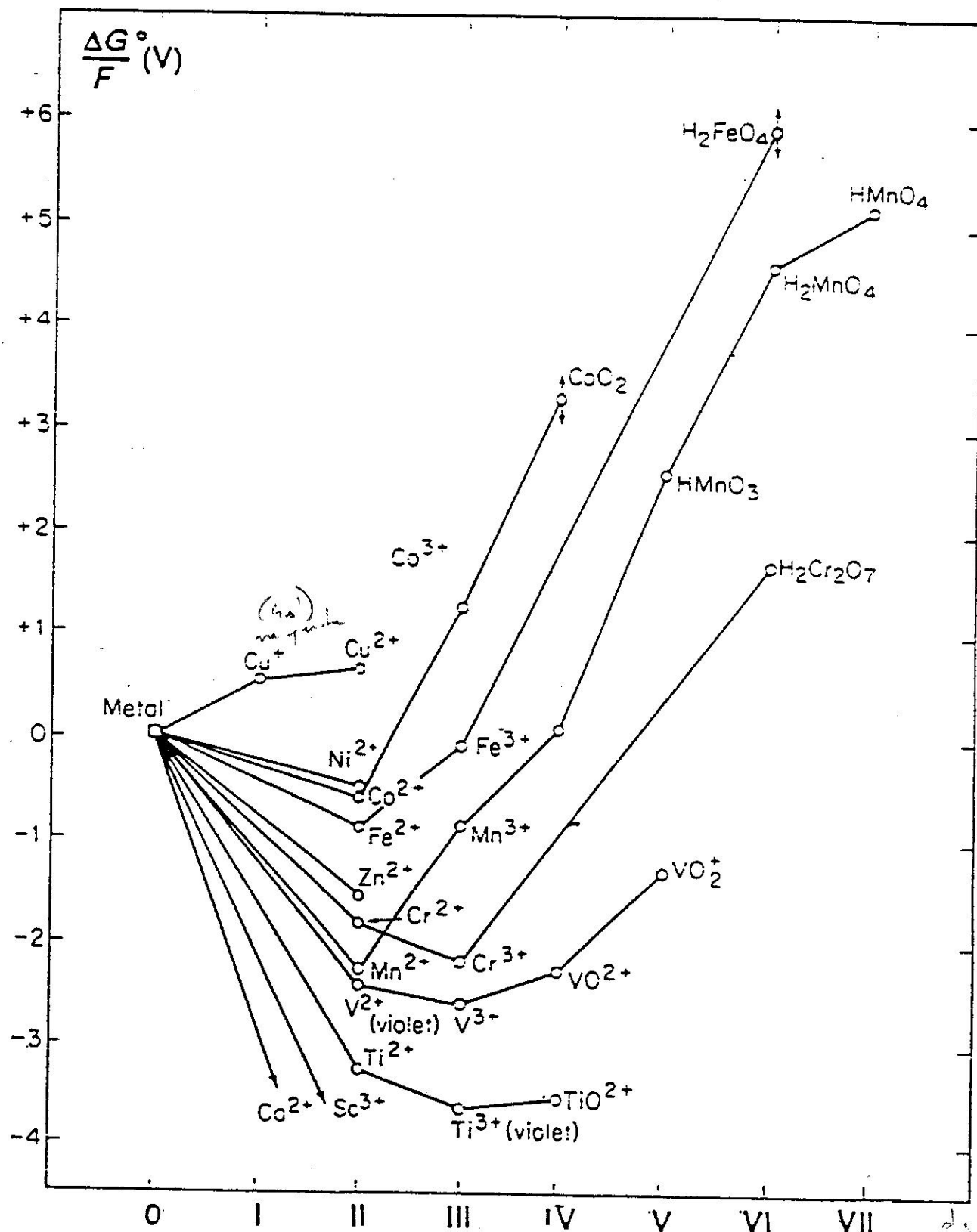


Diagramme.
 Potentiel pH.
 Brome.
 (25° satg/l)



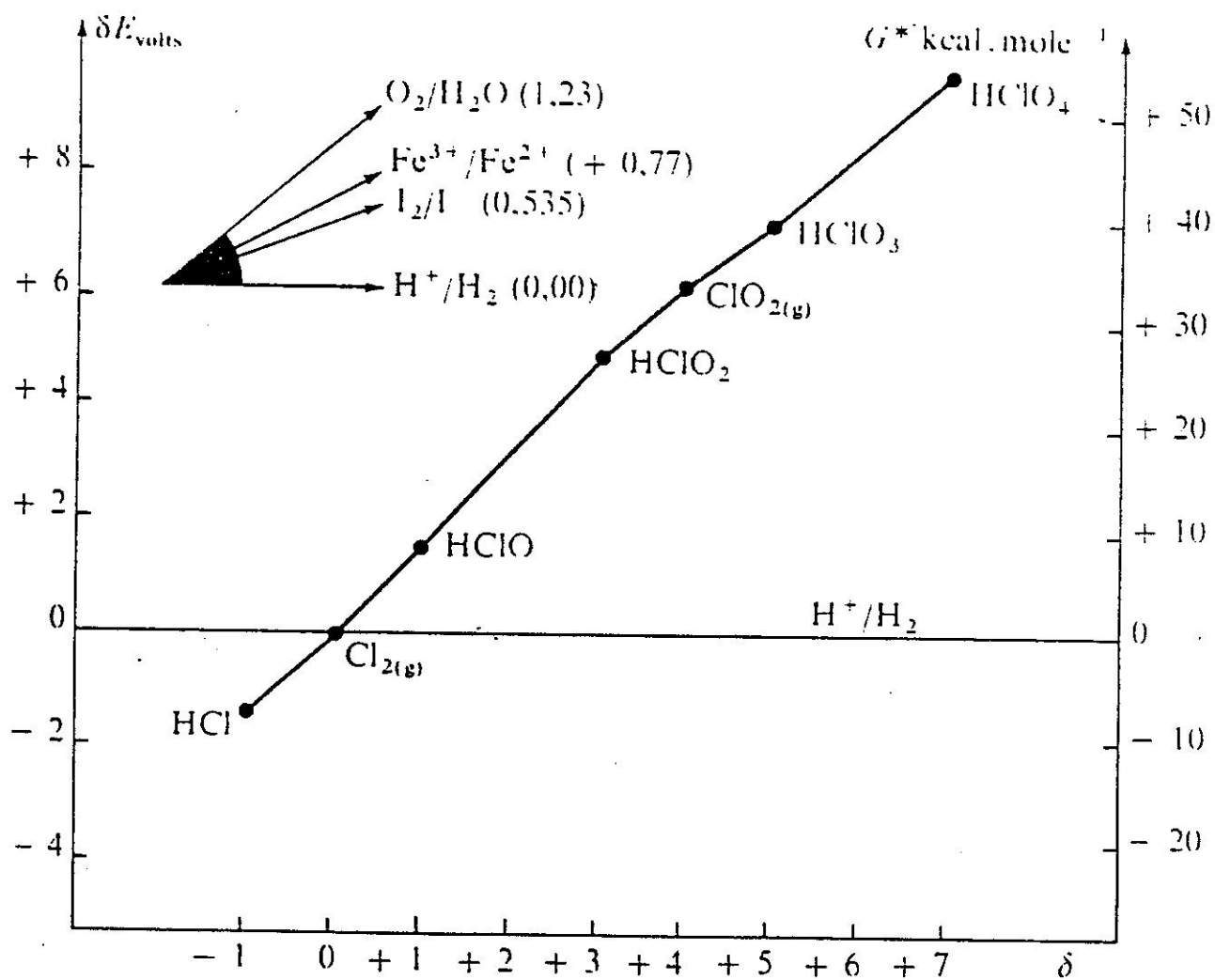






(IIA) d_1
 Mg Sc Ti V Cr Mn Fe Co
 Ca

d_9 d_{10}
 Cu Zn
 4d (IB)
 5d



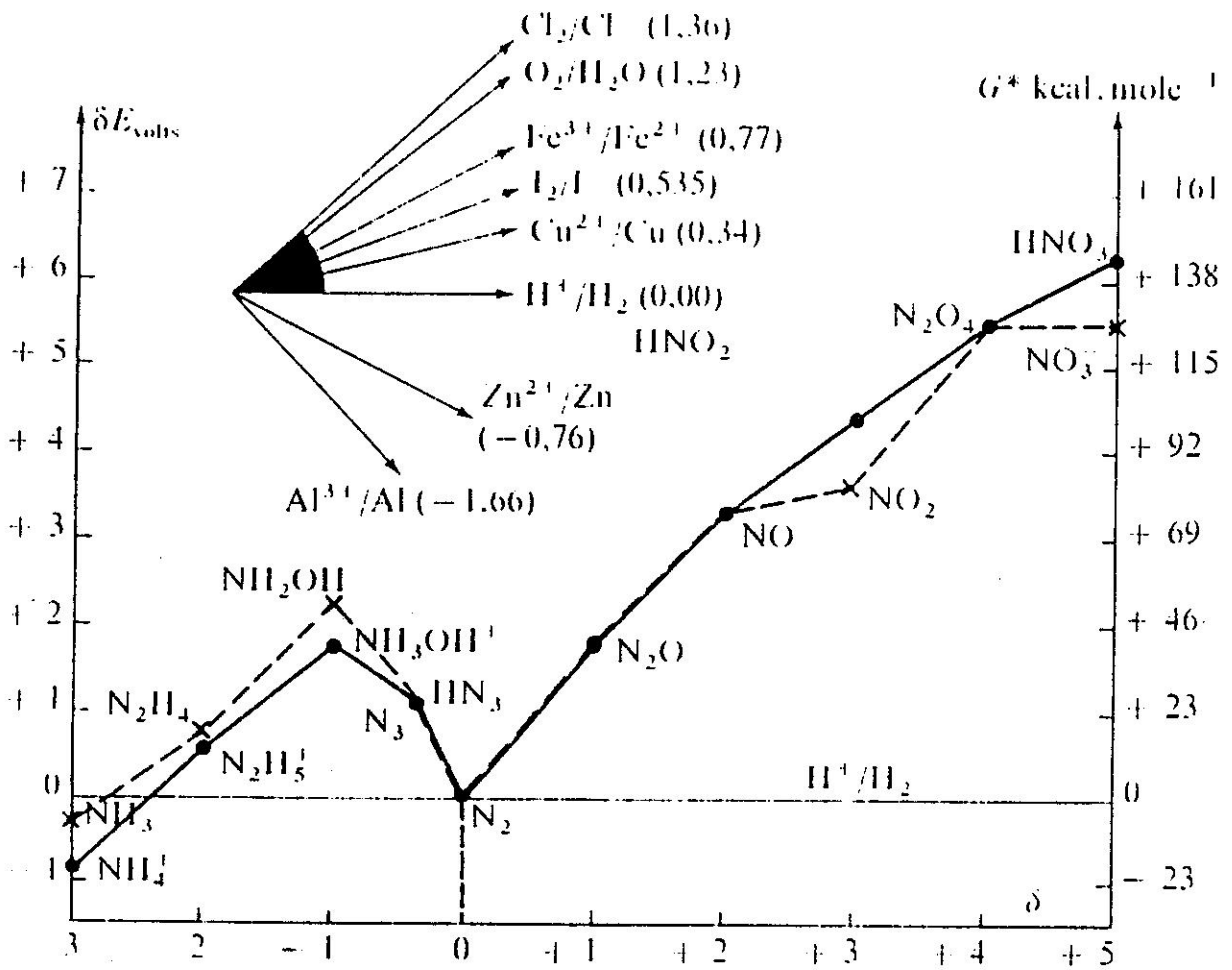


Diagramme rédox de l'azote (b)
 trait plein (pH = 0), trait pointillé (pH = 14)

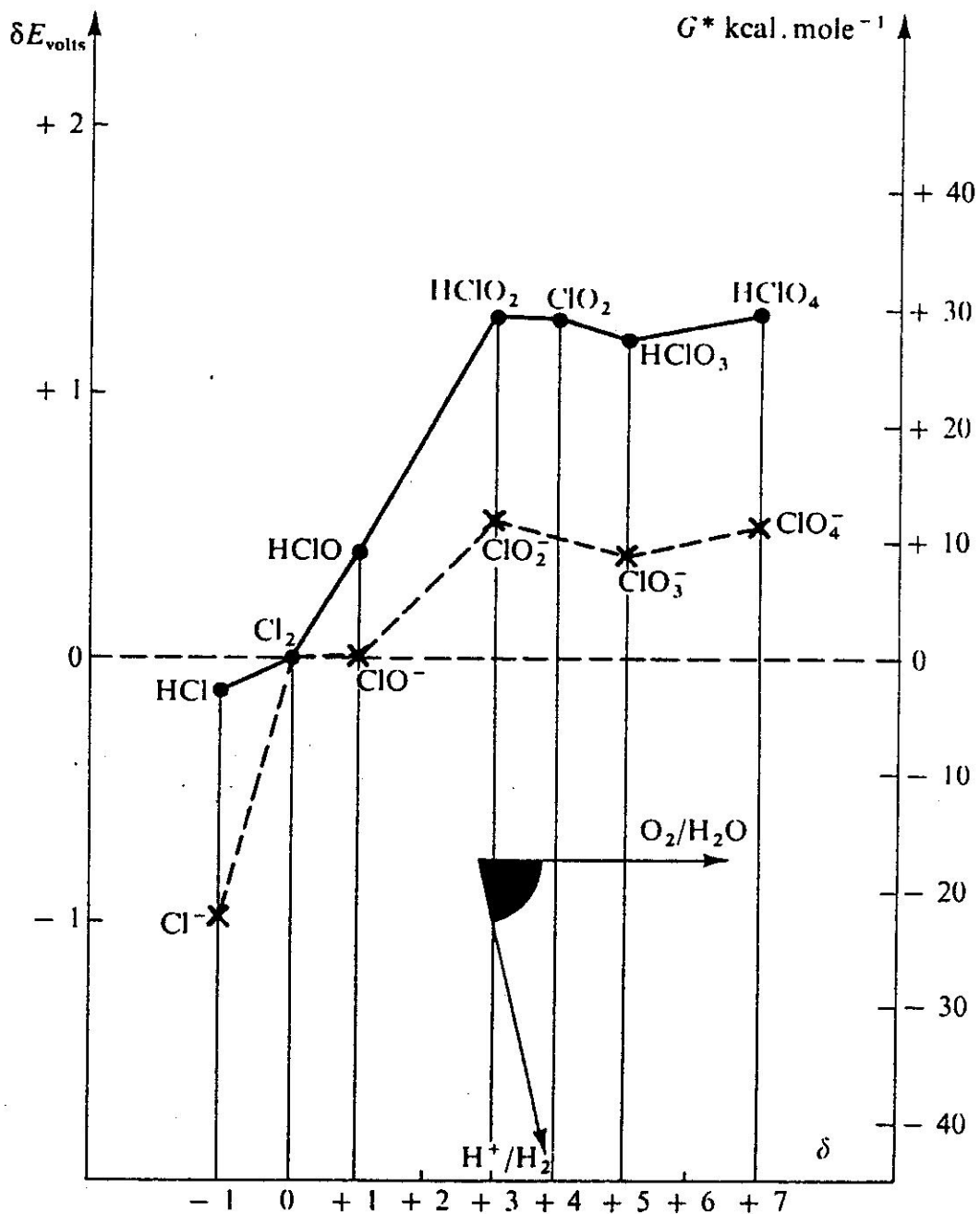
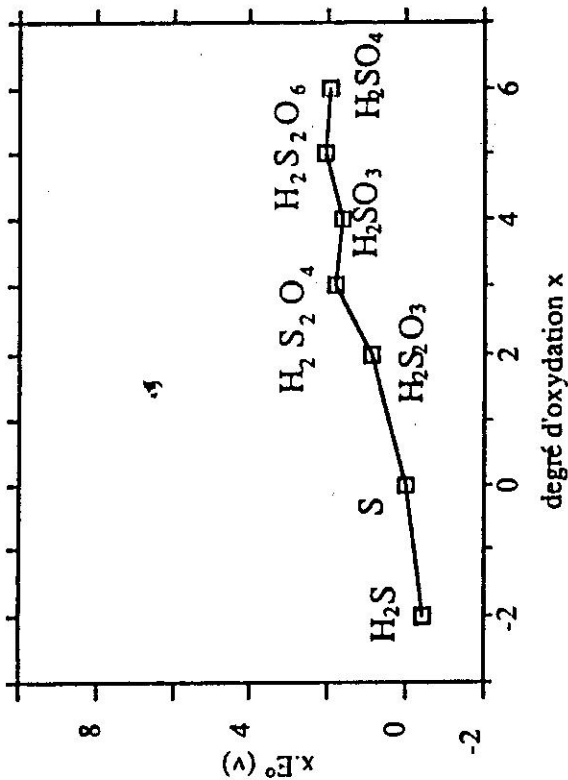
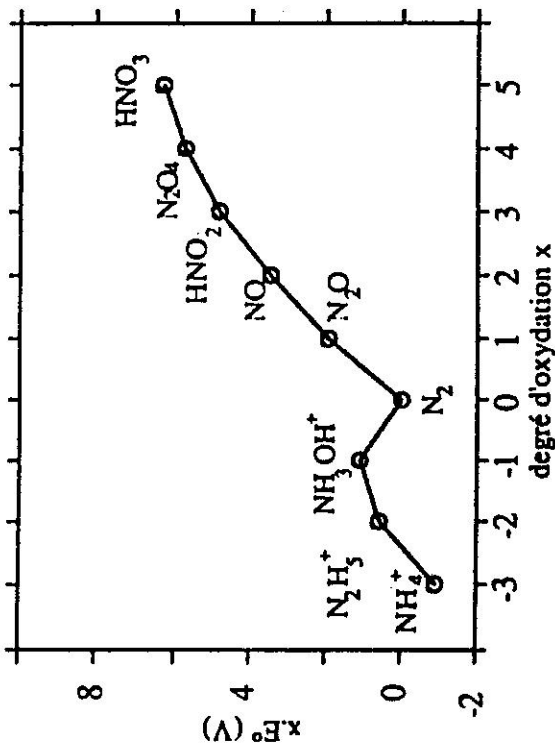
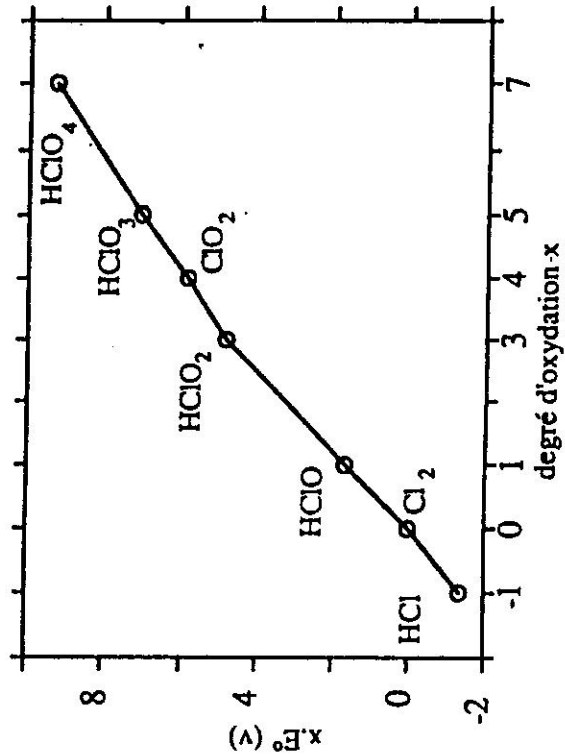
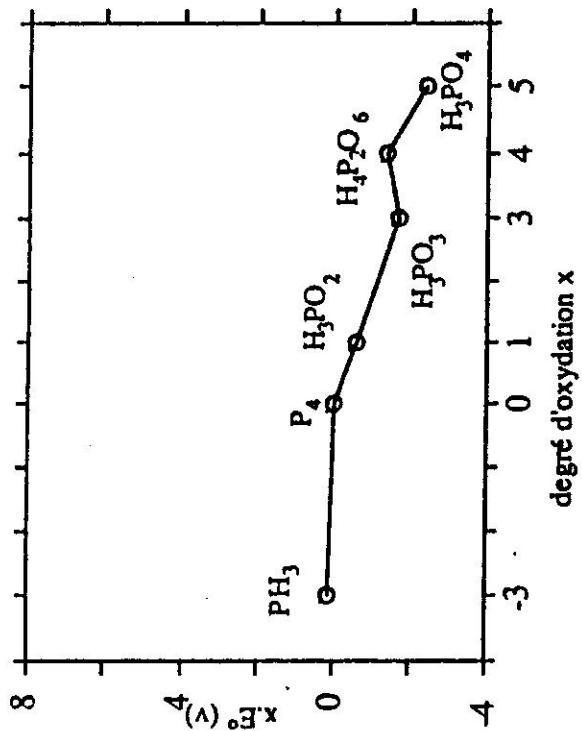
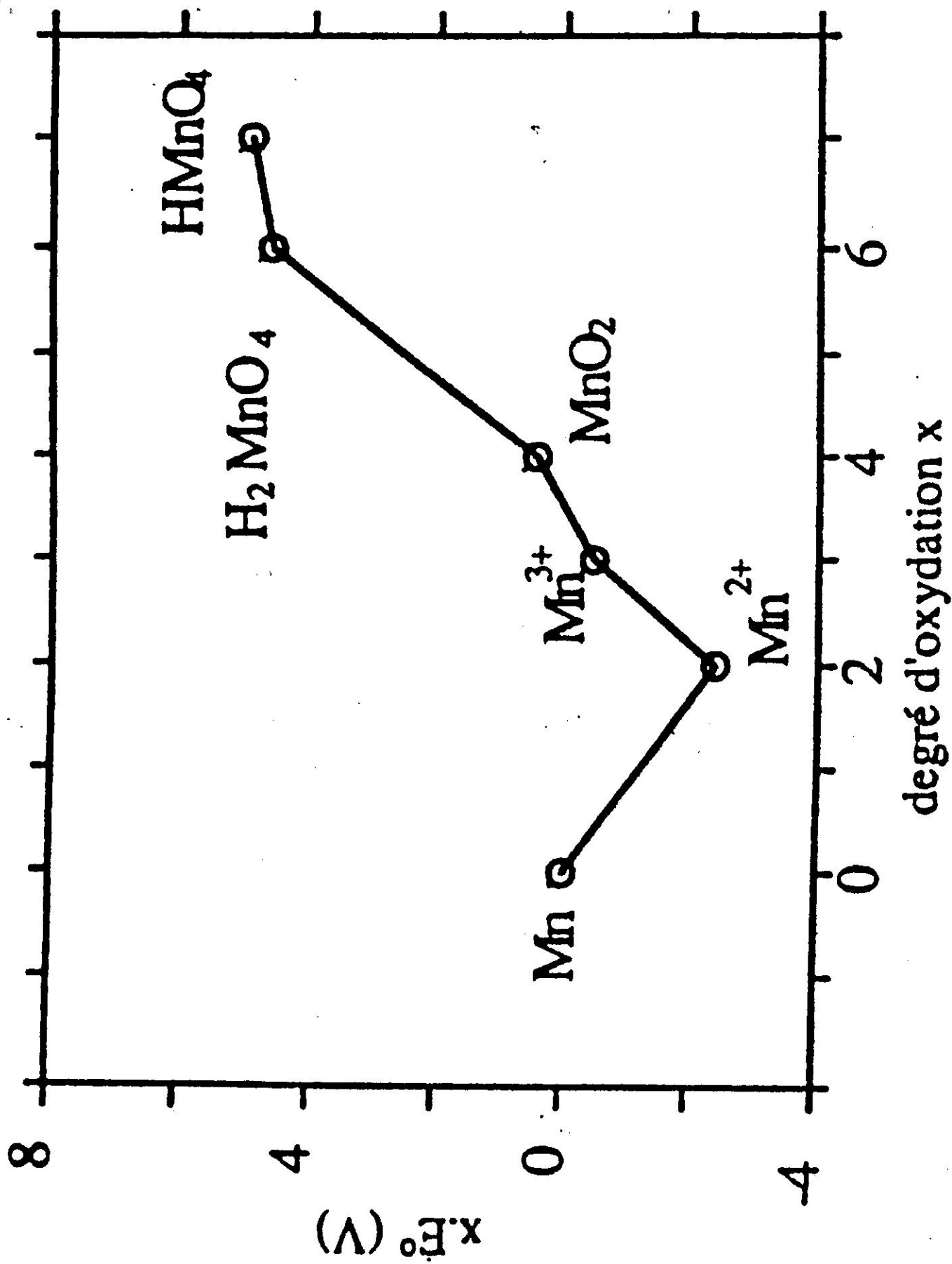


Diagramme rédox du chlore à 25 °C (couple de référence $\text{O}_2/\text{H}_2\text{O}$)
 trait plein (pH = 0) trait pointillé (pH = 14).





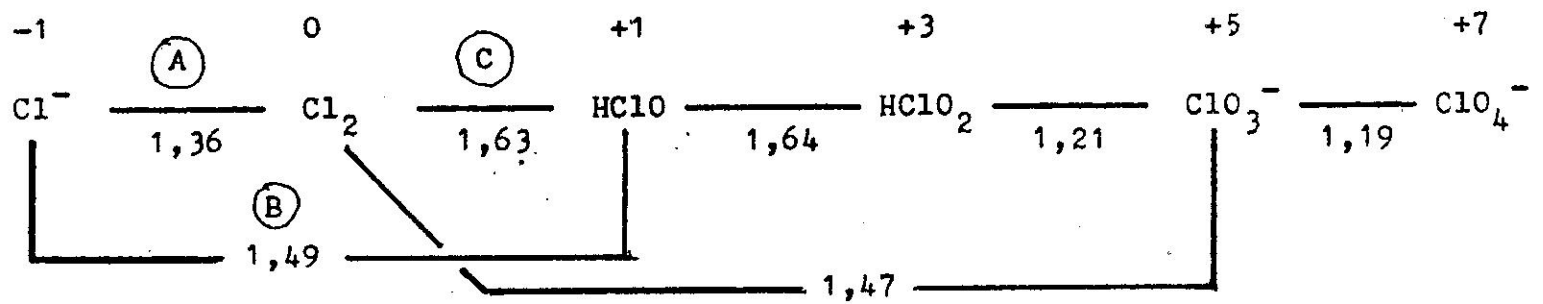
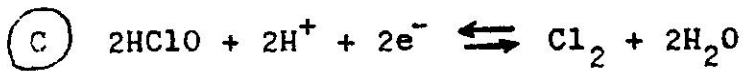
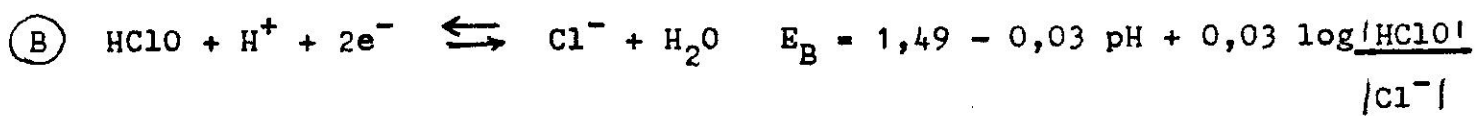
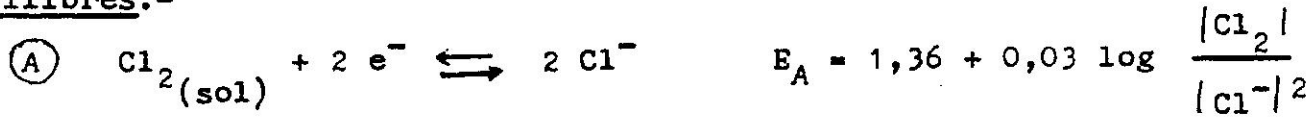
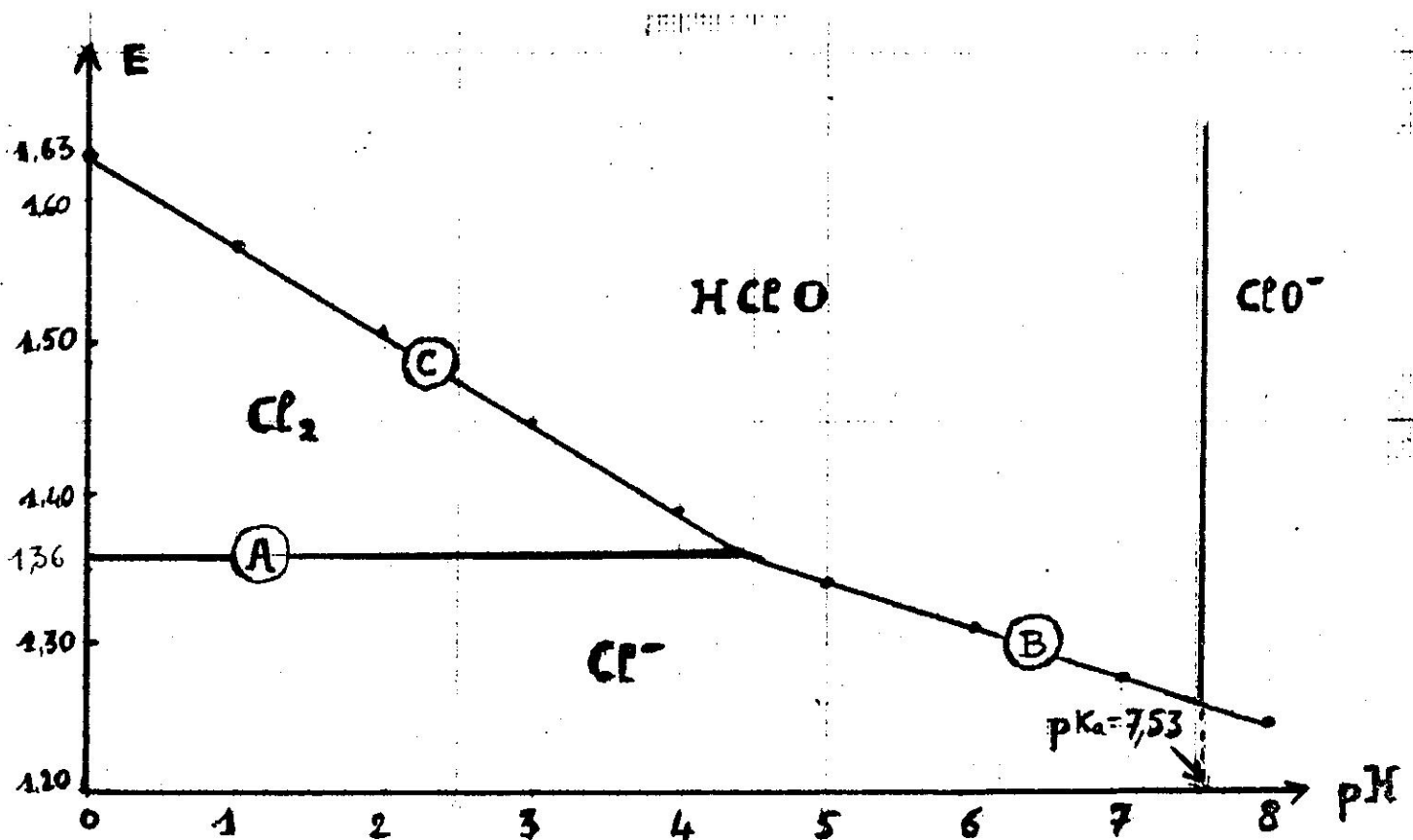


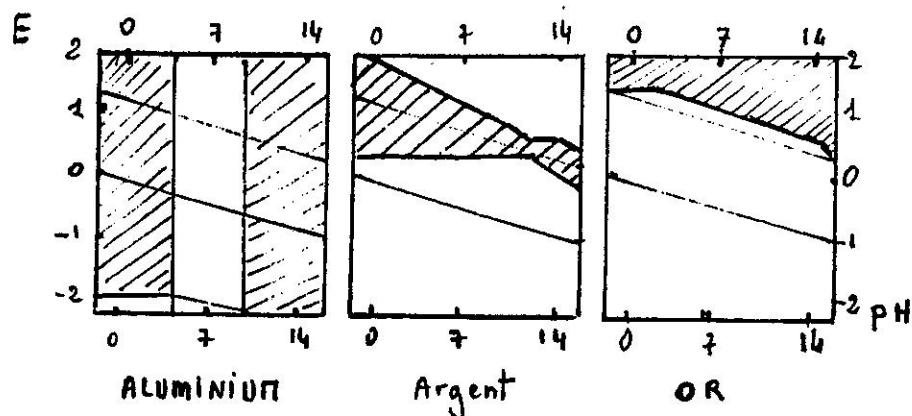
Diagramme de LATIMER à (ph = 0)

Equilibres.-

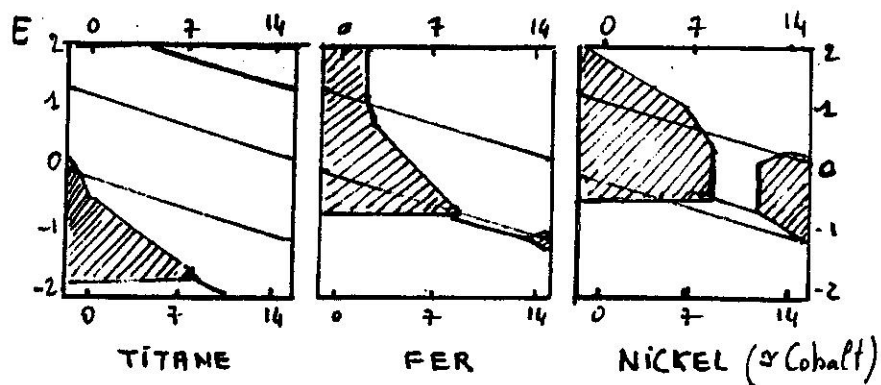
Cet équilibre C résulte de $2(\text{B}) - (\text{A})$ soit $E_C = 2E_B - E_A$

$$\text{d'où } E_C = 1,63 - 0,06 \text{ pH} + 0,03 \log \frac{|\text{HClO}|^2}{|\text{Cl}_2|}$$





- (1) **METAUX RESISTANTS A L'EAU PURE**
 Les zones hachurées indiquent les domaines théoriques de corrosion
 Les zones non hachurées indiquent les domaines théoriques d'immunité et de passivation.



- (2) **METAUX PASSIVABLES ET ACTIVABLES**
 Les zones hachurées indiquent les domaines théoriques de corrosion
 Les zones non hachurées indiquent les domaines théoriques d'immunité et de passivation